

Synthesis of Lauric-Myristic Acid/Activated Carbon Composite as a New Shape-Stabilized Energy Storage Material

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In this work, a new composite phase change material (CPCM) with lauric-myristic acid (LA-MA) eutectic as PCM and activated carbon (AC) was used as supporting material with four different mass ratios of 5.0:5.0, 5.5:4.5, 6.0:4.0, and 6.5:3.5, respectively. The properties and microstructure of LA-MA/AC were analyzed by some characterization methods. The results show that the composite process of LA-MA eutectic and AC was a simple physical mixing and no new chemical bonds were found. The fusion and freeze temperature, enthalpy of the samples were measured by differential scanning calorimetry (DSC), and the residual weight of the samples was analyzed by thermogravimeter (TGA). It was shown that the fusion and freeze temperature of LA-MA eutectic separately were 32.42 °C and 33.63 °C, and its fusion enthalpy and freeze enthalpy were 152.64 J/g and 148.8 J/g, respectively. TGA data shows that the thermal stability of LA-MA eutectic was obviously improved by adding AC as a support material. The results of this study can be available for reference to solar energy storage applications.

Keywords: lauric-myristic eutectic, compound materials, activated carbon, phase change materials.

1. INTRODUCTION

The concept of latent heat storage is to use materials to absorb or release the enthalpy of phase change during the phase transition, and it has the characteristics of high energy storage density and almost no temperature fluctuation in the process of releasing thermal energy. Among its four phase change forms, solid-liquid phase change materials (SLPCMs) have a relatively higher enthalpy than solid-solid, and the volume change is smaller than solid-gas and gas-liquid PCM. Therefore, it has been widely concerned in recent years.

SLPCMs can be divided into two types: inorganic and organic materials, according to the composition of different chemical elements. Inorganic materials have the advantage of high thermal conductivity, good flame retardancy and high volumetric energy storage density [1]. It has obvious drawbacks that cannot be ignored, such as supercooling, phase separation and violent corrosivity, which limit the energy conversion in time. Organic materials have the advantage of low toxicity, low corrosion, economy and no phase separation compare to inorganic materials, and it can be used for energy storage systems with low melting temperature (< 150 °C) [1]. Organic materials have been designed for textiles [2], cold storage [3], building materials [4, 5], solar energy storage [6, 7], lithium batteries [8, 9], waste heat recovery [10, 11], air conditioning [12], etc. At the same time, organic materials (paraffin, fatty acids, eutectic materials, etc.) have the phenomenon of leakage and low heat exchange coefficient in the process of phase transformation. In the past, researchers had developed many methods to combine it with support materials to manufacture shape-stabilized composite PCM (CPCM) to overcome these shortcomings.

At present, several supporting materials that have received much attention include expanded graphite (EG) [3, 13–15], carbon nanotubes (CNT) [16, 17], expanded perlite (EP) [18, 19], activated carbon (AC) [20–22], expanded vermiculite (EVT) [23] and graphite nanoplate (Gr) [24]. Su et al [25] studied n-octadecane/stearic acid (SA-ODE) eutectic mixture as PCM and hexagonal boron nitride (HBN) was served as support material to manufacture CPCM. It was concluded that the optimal ratio was the SA-ODE eutectic with 10 % HBN, and the heat exchange coefficient is increased by 8.0 % when the temperature was 20 °C. Kim et al. [14] studied the octadecane (OD)/WEPG composite PCM prepared by using the kneader mixing technology to absorb OD as an energy storage material into the support material WEPG; the results indicate that when the WEPG content in the OD/WEPG composite exceeds 30 %, it keeps the form stable during the temperature rises from 23 to 80 °C. Ince et al. [24] prepared a myristic acid (MA)/Gr composite material using MA as a thermal energy conversion carrier and Gr was served as a matrix material; their finding when the content of Gr in composite PCM at 2 %, the thermal conductivity increased 38 %. Hu et al. [20] investigated 1-Hexadecanol-Palmitic acid (HD-PA)/AC composite developed by vacuum impregnation; the conclusion was that the heat exchange coefficient and thermal stability of the composites were enhanced. Putra et al. [16] synthesized beeswax/ multi-wall CNTs composite PCM by the impregnation method; they found that adding multi-wall CNTs can improve the thermal performance and stability of solid and liquid beeswax.

In this work, the eutectic mixture of lauric-myristic acid (LA-MA) is selected as PCM due to relatively high enthalpy and low price, AC is used as support material with abundant

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micro-pore structures, huge specific surface area, low density and good thermal conductivity [20]. Four CPCMs were fabricated with the ratio of LA-MA eutectic and AC at 5:5, 5.5:4.5, 6:4, and 6.5:3.5, respectively. The properties and microstructures of LA-MA/AC were analyzed by some characterization methods, and the compositions of LA-MA and AC was verified by the analysis of the chemical structure.

2. EXPERIMENTAL

2.1. Materials

Lauric acid ($C_{12}H_{24}O_2$, AR, pure > 99 %), activated carbon (AC, > 200 mesh), and myristic acid ($C_{14}H_{28}O_2$, AR, pure \geq 99 %) were obtained from Shanghai Aladdin Bio-chem Technology Co., Ltd. Eutectic of lauric-myristic acid (LA-MA) was used as a heat storage material.

2.2. Synthesis of LA-MA eutectic

According to the eutectic theory, LA and MA were blended to form a eutectic material. The eutectic point of LA-MA can be estimated by Schroder's theoretical equation [26, 27]:

$$T = \frac{1}{\left(\frac{1}{T_i} - \frac{R \ln x_i}{\Delta H_i}\right)}, \quad (1)$$

where T represent melting point of eutectic, R stands for gas constant. T_i , x_i and ΔH_i separately represent melting temperature, molar fraction and enthalpy of component i .

Based on formula (1), the molar ratio of LA and MA was calculated at 6.51:3.49. To facilitate mass weighing, the molar ratio was converted into mass ratio, which was determined at 62:38. The mixture that weights 10 g with mass ratio 62:38 for LA and MA is poured into a 500 ml beaker, which placed in a thermostatic magnetic stirring water bath at 70 °C and heated until the mixture completely melted. Then, the mixture was stirred at 600 rpm/min for 1 h to guarantee the mixing uniformity with the constant temperature.

2.3. Preparation of LA-MA/AC composite

In this work, four kinds of CPCMs with different composite ratios were prepared by using AC as the matrix material and LA-MA eutectic as PCM. The specific composition ratio of CPCMs are listed in Table 1. Before mixing, the weighted AC was placed in an air-circulation oven at 70 °C for 16 h to remove the moisture from the porous matrix materials. Then, the dried porous material was placed in a molten eutectic mixture, during which the temperature was maintained at 70 °C, and the rotation speed was maintained at 600 rpm/min and stirred uniformly for 1 hour to a homogeneous mixture. Finally, The CPCMs were placed in a dry oven at 33 °C for 48 h, so that the water evaporates and obtained dried CPCMs.

In order to obtain a shape-stable composite with a high phase change enthalpy, CPCMs were put in an air-circulation oven at 50 °C and heated for 1 h. In Fig. 1, it was found that the weight percentage of LA-MA eutectic in sample was more than 55 %, which was caused by liquid phase leakage.

Table 1. The mass ratio of different CPCM

Samples	LA, g	MA, g	AC, g	Mass ratio of LA-MA/AC
CPCM1	6.20	3.80	10.00	50:50
CPCM2	6.20	3.80	8.20	55:45
CPCM3	6.20	3.80	6.70	60:40
CPCM4	6.20	3.80	5.40	65:35

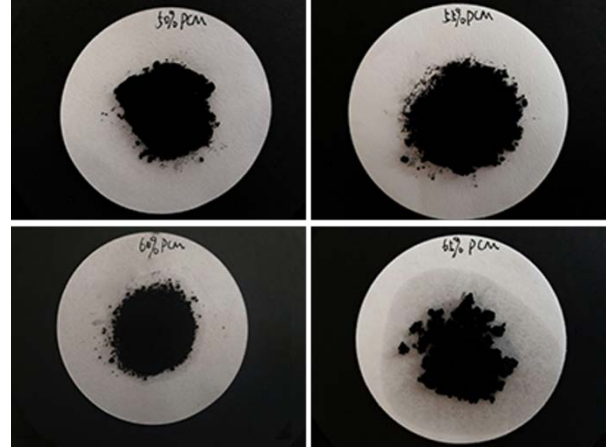


Fig. 1. The leak test of CPCMs with different LA-MA content

2.4. Characterization and properties of CPCMs, AC and LA-MA eutectic

The chemical structures of CPCMs, LA-MA eutectic and AC were determined by Fourier transformation infrared spectrometer (FT-IR, Nicolet5700, spectra range from 4000 ~ 400 cm^{-1} , resolution using KBr pellets). All CPCM samples and AC were treated by gold spraying under vacuum to observe their micro-structure by scanning electron microscope (SEM, JSM6701F). The crystalline phase of CPCMs, LA-MA eutectic, and AC were analyzed by X-ray diffractometer (XRD, D8) within angle range of 10–60° with a speed of 2 θ (5 °/min). The enthalpy and melting-solidifying temperature of the LA-MA/AC were analyzed by differential scanning calorimetry (DSC, DSC 8500) within the temperature range of –10–90 °C at the temperature change of 5 °C/min under a continuous nitrogen flow. The residual weight of the CPCMs was tested by thermogravimetric analysis (TGA, TGA 4000) at room temperature to 700 °C under the condition of continuous nitrogen flow and a heating rise of 20 °C/min.

3. RESULTS AND DISCUSSION

3.1. Analysis chemical structure by FT-IR spectrum

Fig. 2 shows the FT-IR spectra of LA-MA eutectic, AC, and CPCM1-CPCM4. The LA-MA eutectic spectrum shows that there are several main bands at 2917 cm^{-1} , 2845 cm^{-1} , 1697 cm^{-1} , 1471 cm^{-1} , 1302 cm^{-1} , 983 cm^{-1} , and 716 cm^{-1} . The characteristic spectral bands corresponding to the symmetric and antisymmetric stretching vibration of $-CH_3$ and $-CH_2$ chemical bonds appear at 2917 cm^{-1} and 2845 cm^{-1} [17, 28]. The absorption spectral band at 1697 cm^{-1} is produced by stretching vibration of C=O bonds. The spectral band at 1471 cm^{-1} represents the deformation vibration of $-CH$ in $-CH_2$ [29]. The in-plane

deformation vibration and out of plane deformation vibration of $-OH$ functional group results in the characteristic spectral bands at 1302 cm^{-1} and 983 cm^{-1} [30]. The spectral band at 716 cm^{-1} is caused by the swinging vibration of $-CH_2$ group. The characteristic spectral band at 1583 cm^{-1} corresponds to the stretching vibration of $C=C$ bonds. The FT-IR spectra of CPCM1-4 have all absorption spectral bands of binary eutectic acid and the band of AC at 1583 cm^{-1} . The results show that the composite process of LA-MA eutectic and AC was a simple physical mixing and no new chemical bonds were found.

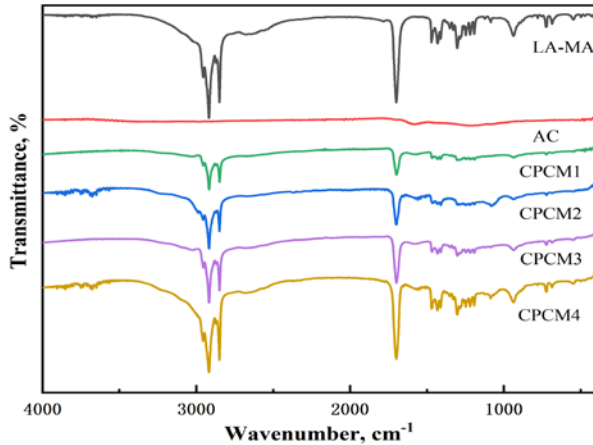


Fig. 2. FT-IR spectra of LA-MA eutectic, AC, CPCM1, CPCM2, CPCM3 and CPCM4

3.2. Analysis of LA-MA/AC by XRD patterns

The XRD curves of LA-MA eutectic mixture, AC, CPCM1, CPCM2, CPCM3 and CPCM4 are shown in Fig. 3.

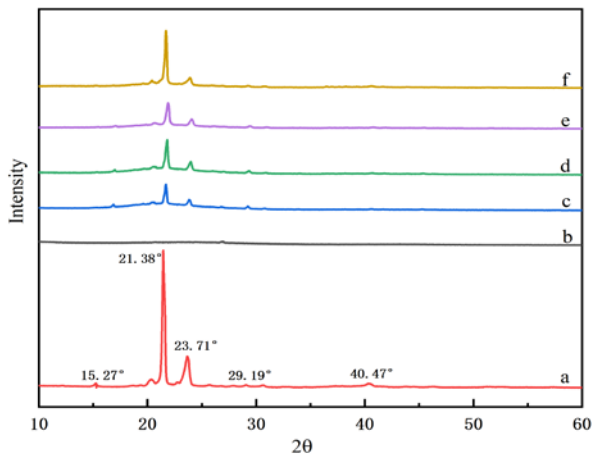


Fig. 3. XRD curves: a–LA-MA eutectic; b–AC; c–CPCM1; d–CPCM2; e–CPCM3; f–CPCM4

The Fig. 3 a presents the XRD pattern of LA-MA eutectic. The curves show that the five intensive peaks of LA-MA eutectic at 15.27° , 21.38° , 23.71° , 29.19° , 40.47° , respectively. The XRD pattern of AC is shown in Fig. 3 b. It indicates that AC has no obvious crystal peaks, so AC is an amorphous structure. Fig. 3 c–f presents the XRD curves of CPCM1, CPCM2, CPCM3 and CPCM4. It can be seen that the four composite PCMs still retain the crystal structure at the same position. It was concluded that the crystal structure of LA-MA eutectic remains during the synthesis, and no new atomic groups are generated.

3.3. Micromorphology analysis of the LA-MA/AC

SEM graphs of morphology of LA-MA eutectic, AC, CPCM1, CPCM2, CPCM3 and CPCM4 are shown in Fig. 4. Fig. 4 a and Fig. 4 f separately present the microstructure image of AC and LA-MA eutectic. It can be observed that the AC has an abundant multi-layered structure and highly porous structure. The comparison between Fig. 4 b, c and Fig. 4 a shows that the AC retains the layered structure and its surface becomes smoother, which shows that the eutectic acid is well absorbed into pore structure due to the effect of surface interaction forces and pore forces. Comparing Fig. 4 d, e and Fig. 4 b, c, it is found that c has completely covered the pore structure of the AC and overflowed on the surface. The comparative analysis shows that the content of LA-MA eutectic in the composite is 55 %, which is the most stable composite.

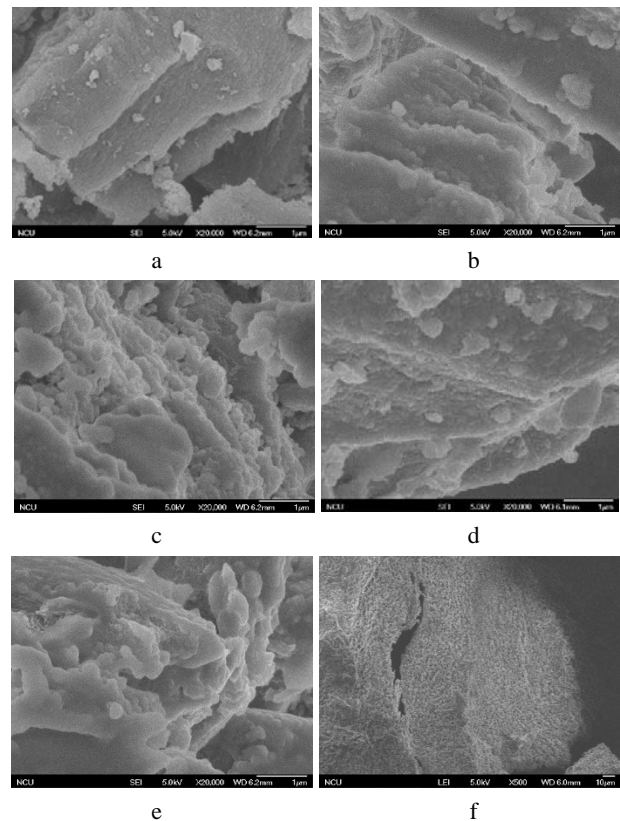


Fig. 4. SEM graphs: a – AC; b – CPCM1; c – CPCM2; d – CPCM3; e – CPCM4; f – LA-MA eutectic.

3.4. The DSC data analysis of CPCMs

The DSC results of the LA-MA eutectic and LA-MA/AC composites are shown in Fig. 5 and Fig. 6. The onset temperature, the peak temperature, enthalpy of fusion and freeze enthalpy are shown in Table 2. It is found in Table 2 that the fusion and freeze temperatures from CPCM1 to CPCM4 reduce as the addition of AC decreases. It may be that the force of LA-MA is enhanced with the load of AC increasing, which impacts on the phase transition process further. The supercooling degree of CPCM1-4 is $0.37\text{ }^\circ\text{C}$, $0.36\text{ }^\circ\text{C}$, $0.32\text{ }^\circ\text{C}$, and $0.27\text{ }^\circ\text{C}$, respectively. The reasons are cause by the AC can prevent the perfect crystallization of pure LA-MA.

Fig. 5 and Fig. 6 illustrate that LA-MA eutectic and CPCM samples only have a peak, indicating that only solid-liquid phase transition during the phase change.

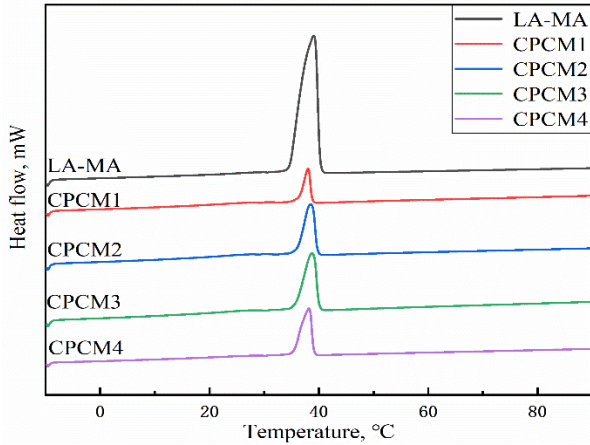


Fig. 5. Melting DSC curves of LA-MA eutectic and LA-MA/AC composites

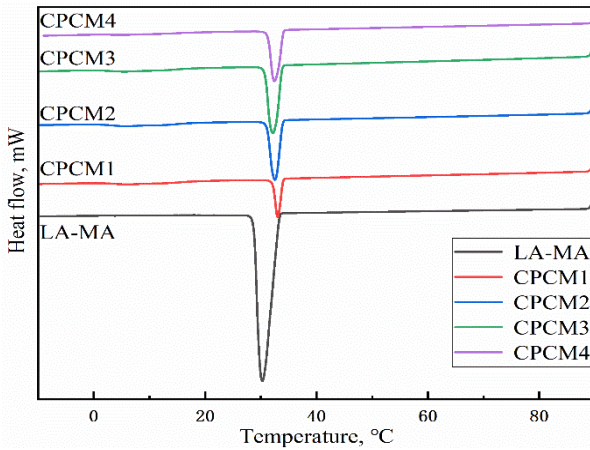


Fig. 6. Solidifying DSC curves of LA-MA eutectic and LA-MA/AC composites

The phenomenon indicates that the addition of AC did not change the energy exchange mode of PCM. The enthalpy of CPCM in Table 2 verifies that the enthalpy increases with the proportion of LA-MA eutectic increases. Therefore, only the LA-MA eutectic in the composites

Table 2. DSC data of the LA-MA eutectic and CPCM samples

Samples	Solid to liquid			Liquid to Solid		
	Onset temperature, °C	Peak temperature, °C	Enthalpy, J·g ⁻¹	Onset temperature, °C	Peak temperature, °C	Enthalpy, J·g ⁻¹
LA-MA eutectic	32.42	38.90	152.64	33.63	30.29	148.8
CPCM1	34.89	38.13	21.91	34.52	32.91	19.61
CPCM2	34.77	38.49	35.46	34.41	32.66	32.71
CPCM3	34.64	38.74	47.96	34.32	32.15	45.45
CPCM4	34.52	38.16	61.67	34.25	32.54	56.82

Table 3. Comparison of this study with other literature studies

CPCMs	Fusion point, °C	Fusion enthalpy, J/g	Mass fraction of PCM, %	Reference
Octanoic-lauric acid/expanded graphite	3.6	132.8	93	[3]
Beeswax/multi-walled carbon nanotubes	59.8	91.6	80	[16]
Myristic acid/graphite nanoplates	54.0	181	99	[18]
Capric-palmitic-stearic acid/ activated carbon	19.82	34.62	55	[21]
LA-MA/AC	38.49	35.46	55	Present study

undergoes a phase transition to convert thermal energy. Based on DSC data, the best composite ratio can be selected. It is found that the melting enthalpy (61.67 J/g) and solidification enthalpy (56.82 J/g) of CPCM4 are higher than those of the other three. However, when the weight proportion of LA-MA in the composite was more than 55 %, there was liquid phase leakage. Therefore, the CPCM2 is the best composite with melting and freezing temperature are 34.77 °C and 34.41 °C, respectively. And its melting enthalpy and freezing enthalpy separately are 35.46 J/g and 32.71 J/g.

The theoretical enthalpy of samples can be estimated by Eq. 2:

$$\Delta H_{LA-MA/AC} = \eta \cdot \Delta H_{LA-MA}, \quad (2)$$

where $\Delta H_{LA-MA/AC}$, η and ΔH_{LA-MA} separately represent the theoretical enthalpy of CPCMs (J/g), mass content of PCM in CPCM and pure LA-MA eutectic latent heat (J/g). It is calculated that the theoretical value is quite different from the DSC test value. This result may be due to the interaction force between the PCM is strengthened so that the part of PCM cannot crystallize freely.

The performance parameters compared with previous experimental studies using fatty acids as energy storage carrier are listed in Table 3, including fusion point, fusion enthalpy and the mass fraction of fatty acids in the CPCM. It can be found that the LA-MA/AC prepared in this study has a good heat storage capacity under the same support material.

3.5. TGA data analysis of LA-MA/AC composites

The TGA and derivative thermogravimetric (DT) curves of LA-MA/AC composites are shown in Fig. 7 and Fig. 8, respectively. The data of TGA and DT of LA-MA/AC composites are shown in Table 4, which contains the onset decomposition temperature (T_{onset}), the temperature of maximum mass loss rate (T_{max}), the final decomposition temperature (T_{end}), the maximum weight loss rate (W_{max}) and the residual weight at 700 °C. The Fig. 7 and Table 4 indicates that the LA-MA eutectic begins to decomposition at 148.12 °C and reaches the maximum decomposition rate of 46.34 %/min at 265.33 °C.

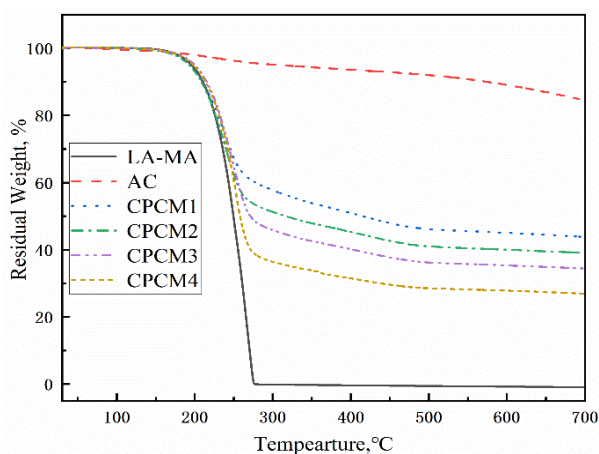


Fig. 7. TGA curves of LA-MA eutectic, CPCM1, CPCM2, CPCM3, CPCM4 and AC.

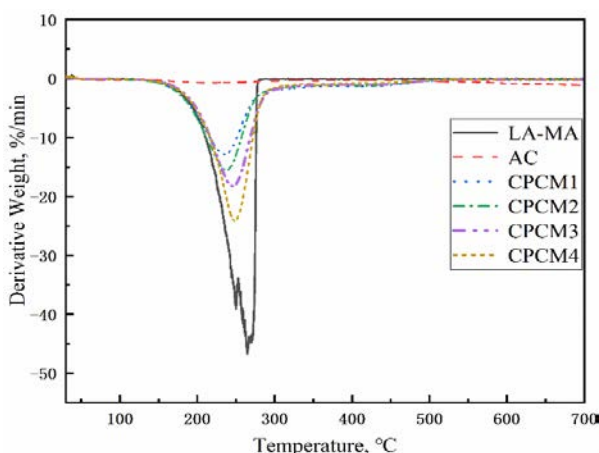


Fig. 8. DT curves of LA-MA eutectic, CPCM1, CPCM2, CPCM3, CPCM4 and AC

CPCM1 to CPCM4 have a similar process of decomposition, the initial decomposition temperature at 134.74 °C, 142.87 °C, 140.46 °C, and 135.57 °C, respectively and the maximum decomposition temperature at 235.9 °C, 240.97 °C, 246.02 °C, and 249.39 °C, respectively. This decomposition temperature is very similar to the decomposition temperature of the LA-MA eutectic, indicating that the main decomposition product is LA-MA eutectic at this temperature. The residual weight of CPCM1, CPCM2, CPCM3, CPCM4 are shown in Table 4, which are 43.71 %, 38.98 %, 34.34 %, and 26.85 %, respectively. However, the maximum decomposition rate decreases with the proportion of AC increases, indicating that AC enhances the thermal stability of the eutectic material. The mass loss value is slightly different from the ratio of LA-MA eutectic in composite PCM and Fig. 8 illustrates that AC has almost no mass loss during heating, indicating that it has good thermal stability. The results show that the thermal stability of the LA-MA eutectic is enhanced by adding AC.

Table 4. TGA data of the LA-MA eutectic, CPCM1, CPCM2, CPCM3, CPCM4 and AC

Samples	T_{onset} , °C	T_{max} , °C	T_{end} , °C	W_{max} , %/min	Residual weight, % at 700 °C
LA-MA eutectic	148.12	265.33	281.11	46.34	0
CPCM1	134.74	235.92	299.65	12.86	43.71
CPCM2	142.87	240.97	304.55	15.38	38.98
CPCM3	140.46	246.02	307.92	18.08	34.34
CPCM4	135.57	249.39	307.16	24.03m	26.85
AC	199.07	–	–	–	84.52

4. CONCLUSIONS

A new composite was prepared with LA-MA eutectic as the PCM and AC was used as the support material, in which the mass fraction of PCM was 50 %, 55 %, 60 % and 65 %, respectively. The FT-IR and XRD results show that the composite process of LA-MA eutectic and AC was a simple physical mixing and no new chemical bonds were found. The DSC data and SEM graphs indicates that the optimal composite ratio of LA-MA eutectic to AC is 5.5:4.5, and the fusion and freeze temperature are 34.77 °C and 34.41 °C, respectively. The difference between fusion temperature and freeze temperature is 0.32 °C. And its melting enthalpy and freezing enthalpy are 35.46 J/g and 32.71 J/g, respectively. This material can be applied for energy storage in solar energy storage system.

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