

Modification of NiFe Layered Double Hydroxide and Photocatalytic Activity

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Modifying materials by ions with multiple valence states is an effective way for enhancing their properties. NiFe layered double hydroxides (NiFe LDHs) were prepared by using a hydrothermal method. Efficient LDH photocatalyst was synthesized by the modification of NiFe LDHs with Fe(III). Electron microscopy, X-ray diffraction, Brunauer-Emmett-Teller surface area, photoluminescence spectroscopy, and UV-Vis diffuse reflectance spectroscopy were used to investigate the morphology, crystal structure, and optical properties of the NiFe LDHs. The photocatalytic and adsorption performances were examined in the degradation of methyl orange. NiFe LDHs exhibited higher adsorption property but lower photocatalytic activity. After modification with Fe(III), the introducing of Fe ions and recrystallization of LDHs led to the distance decreasing between NiFe LDH sheets and exhibit higher photocatalytic performance.

Keywords: layered double hydroxide, modification, photocatalytic activity, adsorption activity.

1. INTRODUCTION

Layered double hydroxides (LDHs) are a class of layered anionic clays, generally expressed by the formula $[M^{II}_{1-x}M^{III}_x(OH)_2](A^{n-})_{x/n} \cdot yH_2O$ [1]. Recently many attentions have been focused on LDHs for their applications in environmental remediation, solar energy storage and conversion, owing to their readily exchanged interlayer anions and tuned composition of the brucite-like layers [2]. However, the process always takes a long time (more than 12 h) [3, 4]. The fast recombination rate of the photogenerated electron-hole pairs and the low removal efficient limit their large-scale practical applications [5, 6]. Therefore, the photodegradation efficiency of LDHs need to further increase.

Modification is one of the common methods for improving their photocatalytic activity. NiFe-LDHs were modified through ion exchanging with $[M(C_2O_4)_3]^{3-}$ ($M = Cr$ and Rh) and exhibited higher magnetic properties [7]. In addition, studies reported that ions with multiple valence states (such as Co , Fe , Cr) can offer an effective pathway for electron-hole transport [8, 9]. For example, the $Co(II)$ in $CoAl$ -LDHs helped to capture $Pd(II)$ species through an in situ redox reaction. The obtained $CoAl$ -LDH/ Pd composites had excellent stability and recyclability [10]. The Cr^{3+} in the Au/Cr -substituted hydroxalite could form the Cr^{3+} - Cr^{6+} redox cycle and make it an efficient catalyst for aerobic alcohol oxidation [11]. From this viewpoint, the multiple valence states in LDHs may provide a new means of capturing and separating photogenerated electron-hole pairs and thus improving the photocatalytic activity of the LDHs. This paper presents a simple hydrothermal method for preparing NiFe LDHs and the modification of NiFe LDHs using $Fe(III)$ -citrate complex under water bath heating. NiFe

LDHs exhibited excellent adsorption property. The adsorption amount for methyl orange reached 968.5 mg/g only within 30 min, which much higher than that of reported values [12–14]. After modification, NiFe LDHs showed better photocatalytic activity and the removal rate reached 95.4 % within 30 min. The modification mechanism was investigated, which provides a facile and effective approach for changing the property of LDHs.

2. EXPERIMENTAL DETAILS

2.1. Synthesis and modification of NiFe LDHs

All chemicals and reagents used in the experiment were analytical grade without further purification.

To prepare NiFe LDHs: 15 mmol of $Ni(NO_3)_2 \cdot 6H_2O$, 5 mmol of $Fe(NO_3)_3 \cdot 9H_2O$, 20 mmol of NH_4F , and 50 mmol of urea were dissolved in 100 mL of deionized water under stirring to form a clear solution. The solution was transferred to an autoclave and maintained at 110 °C for 8 h. After the reaction, the product was separated by centrifugation and washed with deionized water three times. The product was dried at room temperature and labeled as NiFe LDHs. Studies [3, 4] have reported that the initial mole ratio of M^{2+} to M^{3+} was 1 ~ 4. After reaction, the mole ratio of M^{2+} to M^{3+} is usually different with that of initial mole ratio. In our experiments, the mole ratio of Ni/Fe is chosen to be 3:1.

To modify NiFe LDHs: ferric ammonium citrate ($(NH_4)_3[Fe(Cit)_2]$) was dissolved in deionized water to form a clear $Fe(III)$ complex solution (25 mmol/L). NiFe LDHs (1.5 g) was added in the above solution under vigorously stirring to form a uniform suspension. The suspension was heated at 75 °C for 12 h under water bath heating. The following process was the same as that of NiFe LDHs. The product was labeled as NiFe-Fe LDHs.

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2.2. Characterization

X-ray powder diffraction (XRD) patterns were recorded using a D8 ADVANCE X-ray diffractometer with Cu K α radiation ($\lambda = 1.5406 \text{ \AA}$). The scanning electron microscopy (SEM) images were recorded on a FEI-Sirion 200 field emission scanning electron microscope. The transmission electron microscopy (TEM) images, selected-area electron diffraction (SAED) patterns, high-resolution transmission electron microscopy (HRTEM) images, and energy dispersive spectroscopy (EDS) spectra were taken with a FEI-Tecna G2 field emission transmission electron microscope. The X-ray photoelectron spectra (XPS) were measured with an ESCALAB 250Xi X-ray photoelectron spectrometer with Al K α (1486 eV) as the excitation source. The photocatalytic degradation of methyl orange was conducted in an XPA-7 type photochemical reactor (Xujiang electromechanical plant) equipped with a 300 W medium pressure mercury lamp as UV light source (mean wavelength 365 nm). The UV-Vis diffuse reflectance spectra were recorded by Shimadzu UV-3600 spectrophotometer with an integrating sphere. The N $_2$ adsorption/desorption tests were measured by Brunauer-Emmett-Teller (BET) measurements using an NOVA2200e surface area analyzer. Photoluminescence (PL) spectra at 385 nm were investigated using Hitachi F-2500 FL spectrophotometer comprising of a Xe lamp as excitation source. Fourier transform infrared (FTIR) spectroscopy was obtained on a Thermo Nicolet 5700. The Ni and Fe contents of the LDHs were determined by inductively coupled plasma atomic emission spectroscopy (PE Optima 2100DV ICP-AES) after dissolving a weighed amount of sample with an aqueous HCl solution. The carbon content was measured on a LECO CS-444 analyzer.

2.3. Properties measurements

The photocatalytic activity and adsorption properties of LDHs were measured using methyl orange. Reaction suspensions were prepared by adding LDHs (20 mg) in 50 mL of methyl orange solution (20 mg/L). The suspensions were ultrasonically treated for 15 min and then stirred in the dark for 30 min to ensure adsorption-desorption equilibrium. To test the photocatalytic performance of the LDHs, the suspensions were irradiated by UV light. To test the adsorption performance of the LDHs, the suspensions were continuously stirred in the dark. Analytical samples were drawn from the reaction suspensions after various reaction times and analyzed using a UV-Vis spectrophotometer at 467 nm.

3. RESULTS AND DISCUSSION

The crystal structure was confirmed from X-ray diffraction patterns (XRD, Fig. 1 a). The reflections of (003), (006), (012), (110) were indexed to typical LDHs materials (JCPDS File No. 51-0463). The basal spacing of LDHs could be calculated based on the d-spacing of (00L) planes ($[d(003) + 2d(006)]/2$). According to the XRD pattern, the basal spacing value of NiFe LDHs was about 0.766 nm, which coincide well with the values for CO $_3^{2-}$ intercalated LDHs materials [15]. After modification, they still remained the phase of LDHs with a slightly lower

degree of crystallinity and the basal spacing (0.754 nm), which was lower than that of metallic oxalate complex intercalated LDHs (1.214 nm) [15]. The results suggested that after modification the interlayer ions of LDHs was CO $_3^{2-}$ and no Fe complex intercalated in the interlayer, owing to no inert gas protection during the preparation and modification.

XPS spectra were used to determine the oxidation state of elements before and after modification. The peaks at 856.2 eV and 873.7 eV are attributed to the 2p $_{3/2}$ and 2p $_{1/2}$ spin state of Ni(II) for LDHs, respectively [16]. The other two peaks at about 880 eV and 862 eV can be assigned to their corresponding shake-up satellite lines (Fig. 1 b). The XPS lines for Fe 2p indicate two sharp peaks at 726.1 eV and 712.5 eV corresponding to the 2p $_{3/2}$ and 2p $_{1/2}$ spin states of Fe(III) (Fig. 1 c), with a peak splitting of 13 eV [17]. After modification, Ni 2p and Fe 2p slightly shift to lower energy levels within the permitted error (0.2 eV). No peaks for Fe complex (725.0 eV and 711.5 eV) was observed, which also indicate that no Fe complex intercalated into NiFe LDHs. Both samples also show a peak for F1s, which are from NH $_4$ F (Fig. 1 d). The mole ratio of Ni/Fe for NiFe LDHs was 1.82:1. After the modification, the mole ratio of Ni/Fe was 0.89:1. The decrease in the mole ratio of Ni/Fe after the modification also indicates the successful doping of Fe ions. On the basis of the results of elemental analysis measurement, the chemical composition of NiFe LDHs was estimated to be [Ni $_{0.66}$ Fe $_{0.34}$ (OH) $_2$](CO $_3$) $_{0.17}$ ·yH $_2$ O (Ni, 38.6; Fe, 18.5; C, 1.9), which gives a Ni/Fe mole ratio of 1.92. For NiFe-Fe LDHs, Ni/Fe mole ratio was 0.95 and the chemical composition is [Ni $_{0.49}$ Fe $_{0.51}$ (OH) $_2$](CO $_3$) $_{0.255}$ ·yH $_2$ O.

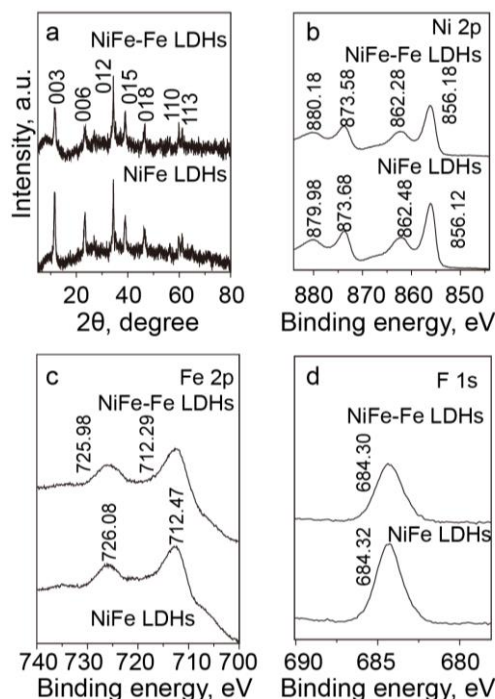


Fig. 1. a – XRD patterns of NiFe LDHs and NiFe-Fe LDHs; b – Ni 2p; c – Fe 2p; d – F1s

FTIR spectra (Fig. 2) of both samples provided evidence for the presence of intercalated CO $_3^{2-}$. Both NiFe

LDHs and NiFe-Fe LDHs showed the strong peak at 1355 cm^{-1} , which can be attributed to the ν_3 vibration of CO_3^{2-} .

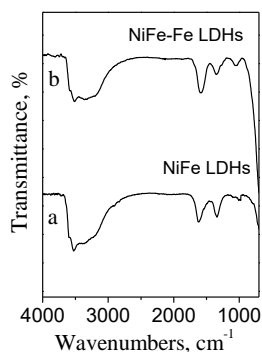


Fig. 2. FTIR spectra: a–NiFe LDHs; b–NiFe-Fe LDHs

SEM image (Fig. 3 a) of NiFe LDHs at different magnifications show NiFe LDHs nanosheets intersecting and aligned vertically to form a flower-like shape under the orienting function of NH_4F [11]. They have a porous structure, and approximately 25 nm thick. The NiFe LDHs nanosheets are transparent to the electron beam, suggesting that they are very thin (Fig. 3 b). The lattice spacing of 0.2658 nm can be indexed to the (012) plane, which matches the reported values of the hexagonal LDH structure well (Fig. 3 c) [18]. After the modification, NiFe-Fe LDHs nanosheets also form the flower-like shape and the nanosheets in the balls are parallel to each other. The distances between these sheets decrease and sheet thickness increases to 200 nm (Fig. 3 d). The lattice spacing of 0.2679 nm is assigned to the (012) plane of NiFe LDHs sheets (Fig 3 e and f). The SAED pattern on a single sheet reveals the crystalline nature of NiFe-Fe LDHs and can be indexed to the (012) reflection of NiFe LDHs (inset in Fig. 3 f). The EDS analysis was employed to determine the composition of NiFe LDHs and NiFe-Fe LDHs, which was similar with the result of XPS analysis. The elemental mapping images also confirm the uniform distribution of Fe and Ni elements in NiFe LDHs and NiFe-Fe LDHs.

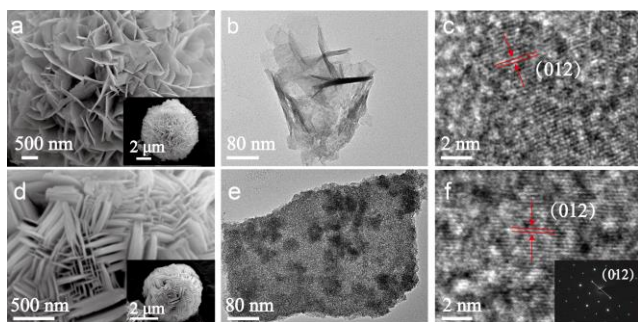


Fig. 3. a, d–SEM images; b, e–TEM images; c, f–HRTEM images of the samples: a–c NiFe LDHs, d–f NiFe-Fe LDHs. The inset of f is corresponding SAED

In deionized water, citrate ions can form soluble chelate complex with Fe^{3+} ($[\text{Fe}(\text{Cit})_2]^{3-}$), which gradually releases Fe^{3+} and retards the hydrolysis rate of Fe^{3+} during the modification process. Therefore, no iron hydroxide or oxide was observed after the modification. $[\text{Fe}(\text{Cit})_2]^{3-}$ was negatively charged and unstable. During the modification,

$[\text{Fe}(\text{Cit})_2]^{3-}$ could absorb on the surface of positive charged NiFe LDHs layers through electrostatic attraction and decompose with longer reaction time [9]. It results in more Fe substituted for Ni in NiFe LDHs and an increasing amount of Fe element, consistent with the preceding XPS and XRD analysis.

The photocatalytic activity in methyl orange photodegradation under UV light irradiation was evaluated (C_0 and C_t are the equilibrium concentrations of methyl orange before and after irradiation, respectively). Before irradiation, NiFe LDHs exhibited better adsorption activity than that of NiFe-Fe LDHs. The adsorption rate reached about 87.65 % for NiFe LDHs and only 1.85 % for NiFe-Fe LDHs after stirring in the dark for 30 min (Fig. 4 a). After 30 min of irradiation, the degradation rate of methyl orange reached about 95.4 % using the NiFe-Fe LDHs as the photocatalyst (Fig. 4 a). The methyl orange solution was decolorized completely after 60 min of irradiation and remained unchanged with a long-time irradiation. When using the NiFe LDHs as the photocatalyst, the photodegradation rate of methyl orange was only 4.85 % although the methyl orange removal rate was about 94 % after 120 min of irradiation (Fig. 4 a). The curve of NiFe LDHs has an upward trend after 30 min of irradiation, owing to the adsorption/desorption equilibrium changing caused by temperature rising under irradiation.

In order to confirm the photocatalytic activity of NiFe-Fe LDHs and NiFe LDHs, blank experiments without irradiation were conducted. A blank experiment using NiFe-Fe LDHs as the photocatalyst without irradiation showed that a small quantity of methyl orange (~ 3.65 %) was adsorbed after 120 min (Fig. 4 a). However, the adsorption rate of methyl orange for NiFe LDHs reached 87.35 % after stirred in the dark for 30 min and 96.85 % continuing to stir for 120 min (Fig. 4 a). According to the preceding analysis, NiFe LDHs had better adsorption activity and NiFe-Fe LDHs exhibited better photocatalytic activity.

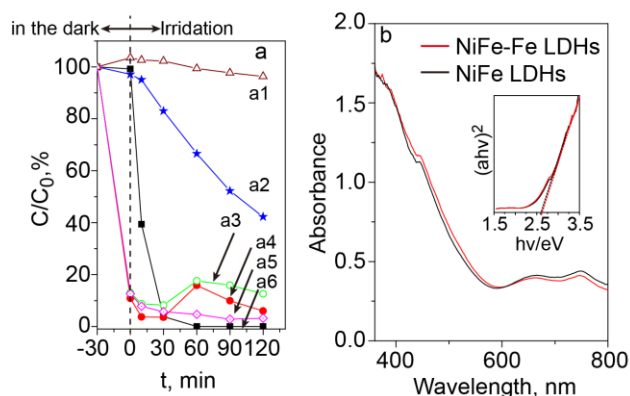


Fig. 4. a–photocatalytic degradation of methyl orange over (a1) blank experiment adding NiFe-Fe LDHs without irradiation, (a2) NiFe LDHs, (a3) blank experiment adding NiFe LDHs without irradiation, and (a4) NiFe-Fe LDHs; b–room temperature UV-Vis diffuse reflectance spectra of NiFe LDHs and NiFe-Fe LDHs. The inset is the plots of $(ah\nu)^2$ vs. photon energy ($h\nu$)

The maximal methyl orange adsorption removal capacity for NiFe LDHs reached 968.5 mg/g, which is much higher than those reported for the anionic dyes

removal (148 mg/g for MgAl-LDHs [12], 350 mg/g for Ni based LDHs [13], and 800 mg/g for ZnNiAl LDO [14]).

In general, the band gap and specific area of materials are related to their photocatalytic activity for the degradation of pollutants. Fig. 4 b shows the room temperature UV-Vis diffuse reflectance spectra of NiFe LDHs and NiFe-Fe LDHs. Both LDHs have absorption in the visible region. The band gap for LDHs was estimated according to the Tauc formula [19] and the plots of $(ah\nu)^2$ versus photon energy ($h\nu$) is shown in the inset of Fig. 4 b. The band gap energies are 2.62 eV for NiFe LDHs and 2.58 eV for NiFe-Fe LDHs.

Fig. 5 a shows the N_2 adsorption/desorption isotherm and pore size distribution curves of NiFe LDHs and NiFe-Fe LDHs. Both LDHs are a type IV isotherm with an H3 hysteresis loop, owing to the aggregation of LDHs nanosheets [20]. A specific surface area of 48.37 m^2/g with a pore size of 17.06 nm is observed for NiFe LDHs, which is a little larger than that of NiFe-Fe LDHs (42.82 m^2/g and 14.27 nm). The pore size is calculated according to the adsorption average pore width ($4V/A$ by BET, where V represents pore volume, A represents surface area).

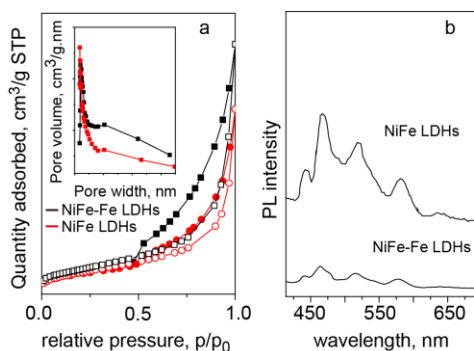


Fig. 5. a— N_2 adsorption/desorption isotherms of NiFe LDHs and NiFe-Fe LDHs, Inset: pore size distribution curves; b—emission spectra at 385 nm for NiFe-Fe LDHs and NiFe LDHs

Compared with NiFe LDHs, NiFe-Fe LDHs exhibited higher photocatalytic activity, although they had similar band gaps and specific area. The mainly differences between them were the mole ratio of Ni to Fe, the thickness of sheet and the distance between sheets. The high recombination rate of photogenerated electron-hole pairs hinders the degradation of methyl orange [21]. The decreasing in distance between NiFe-Fe LDHs sheets could increase particle connectivity and assist in the electron-hole transfer throughout the layered framework, thereby contributing to the high photocatalytic performance [3]. In our experiments, NiFe-Fe LDHs with an appropriate network of pores and more Fe element could facilitate the efficient transportation of photogenerated charges among the LDHs. Studies also reported that the recombination of photogenerated charges can be investigated through photoluminescence (PL) spectroscopy [22]. For band-band PL emission, the lower the PL intensity is, the higher the separation rate of the photogenerated charges. The emission intensity of NiFe-Fe LDHs was lower than that of NiFe LDHs, and hence, a higher photocatalytic activity was observed (Fig. 5 b).

The adsorption of pollutants by LDHs could be ascribed to two modes: the electrostatic interaction between the LDHs layer and charged adsorbates; the ion exchanging between the interlayer ions and adsorbates. NiFe LDHs can be regarded as $Ni(OH)_2$ layered structure, in which parts of Ni^{2+} are replaced by Fe^{3+} , and unequal substitution leads to excess positive charge. The excess positive charges are balanced by intercalation of anions between the layers. The smaller the ratio of Ni to Fe means the more Ni^{2+} replaced by Fe^{3+} and more excessive positive charges in LDHs layers. Meanwhile methyl orange is an anionic organic dye. If electrostatic attraction plays a dominant role, NiFe-Fe LDHs should exhibited better adsorption activity than that of NiFe LDHs. However, NiFe LDHs exhibited higher adsorption activity, which indicated that the adsorption process was mainly through ion exchanging process. NiFe LDHs have larger interlayer distance, specific area and pore size, which lead to higher adsorption property than that of NiFe-Fe LDHs. Meanwhile the larger distance between NiFe LDHs nanosheets is in favor of the diffusion and adsorption of methyl orange [23]. To investigate the stability, NiFe-Fe LDHs renewability experiment was carried out by repeating photocatalytic degradation of methyl orange. After each cycle, NiFe-Fe LDHs was washed three times by water for the next photocatalytic experiment. The degradation ratio of NiFe-Fe LDHs was maintained at 90% at 30 min after four cycle experiments (Fig. 6).

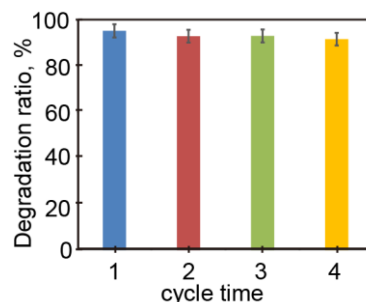


Fig. 6. The regeneration performance of NiFe-Fe LDHs for four cycle experiments

NiFe-Fe LDHs before and after four cycle photocatalytic reactions were examined by XRD, SEM and TEM. As shown in Fig. 7, both the fresh and reacted NiFe-Fe LDHs exhibited a single-phase diffraction pattern assigned to layer structure LDHs, indicating that the original LDHs crystal structure was retained even after four cycles of photocatalytic degradation of methyl orange.

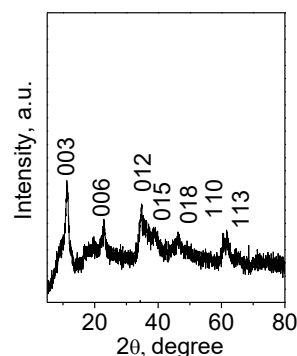


Fig. 7. XRD pattern of NiFe-Fe LDHs after photocatalytic process

And no noticeable change was observed in the particle morphology of NiFe-Fe LDHs before and after photocatalytic reaction (Fig. 8). NiFe-Fe LDHs exhibited good stability and regeneration capacity for the degradation of methyl orange.

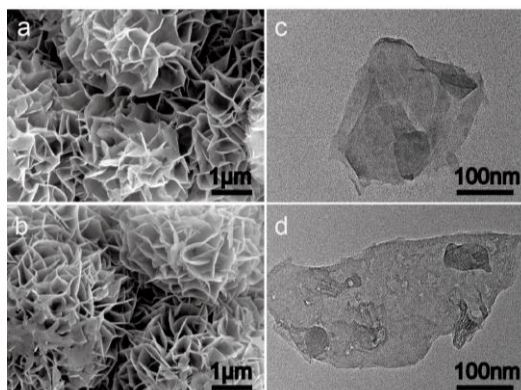


Fig. 8. SEM images of NiFe-Fe LDHs: a – before photocatalytic experiment; b – after four cycle photocatalytic experiment. TEM images of NiFe-Fe LDHs: c – before photocatalytic experiment; d – after four cycle photocatalytic experiment

4. CONCLUSIONS

NiFe LDHs nanosheets with excellent adsorption property were synthesized using a simple hydrothermal method. The larger interlayer distance, specific area and pore size contributed to the higher adsorption activity of NiFe LDHs. After the modification with negatively charged Fe(III) complexes, NiFe LDHs displayed better photocatalytic activity in the degradation of methyl orange, owing to the photogenerated carriers having a longer lifetime for participating in photocatalytic reactions.

Acknowledgments

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